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Erratum

# Erratum to "Primary Li–air cell development" [J. Power Sources 196 (3) (2011) 1498–1502]

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The publisher regrets that in the original printing of the above article Table 1 was printed incorrectly. The correct Table 1 is as follows.

The publisher would like to apologise for any inconvenience this may have caused to the authors of this article/(and) readers of the journal.

**Table 1**
Theoretical specific energy and capacity comparisons for selected systems.

| Metal–air and Li-ion systems (organic or aqueous electrolyte solution) | OCV (V) | Specific energy (Wh/kg) | Specific capacity (mAh/g) |
|---|---|---|---|
| $2Li + ½O_2 \rightarrow Li_2O$ (aprotic organic sln) | 2.913 | 11,248[a] | 3862 |
| $Li + ½O_2 \rightarrow ½Li_2O_2$ (aprotic organic) | 2.959 | 11,425[a] | 3862 |
| $2Li + ½O_2 + H_2SO_4 \Leftrightarrow Li_2SO_4 + H_2O$ (aqueous) | 4.274 | 2046[a] | 479 |
| $2Li + ½O_2 + 2HCl \Leftrightarrow 2LiCl + H_2O$ (aqueous) | 4.274 | 2640[a] | 616 |
| $2Li + ½O_2 + H_2O \Leftrightarrow 2LiOH$ (aqueous) | 3.446 | 5789[a] | 1681 |
| $Zn + ½O_2 \rightarrow ZnO$ (aqueous) | 1.650 | 1353[a] | 820 |
| $x6C + LiCoO_2 \Leftrightarrow xLiC_6 + Li_{1-x}CoO_2$ (organic) | ~4.2 | 420[b] | 139[b] |
| $Li + H_2O$ (in seawater) $\Leftrightarrow LiOH + ½H_2$ | 2.512 | 9701[c] | 3862[c] |

[a] The molecular mass of $O_2$ is not included in these calculations because $O_2$ is freely available from the atmosphere and therefore does not have to be stored in the battery or cell.
[b] Based on $x = 0.5$ in $Li_{1-x}CoO_2$.
[c] The molecular mass of $H_2O$ is not included since it is freely available from seawater (pH 8.2) and does not have to be stored in the battery or cell.

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